

CHEM 109, Lecture 1

Structure – Property Relationships

Acid-Base Chemistry

- McMurry & Begley (M&B) Chapter 1.1-1.2 (posted on CHEM 109 website)
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Revisiting General Chemistry through Structure – Property Relationships

- *Before class:* brainstorm your current understanding of any or all of the terms below
- *During class:* instructions provided for collaborative activity (concept map)
- *Discussion section:* complete a concept map with any one column of terms below

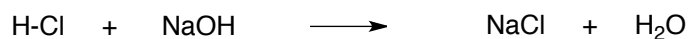
COLUMN 1	COLUMN 2	COLUMN 3	COLUMN 4
Lewis structure	Intramolecular Forces	Boiling	Polar
Skeletal structure	Ionic Bond	Melting	Non-polar
Hybridization	Covalent Bond	Dissolve	Nucleophile
Molecular Geometry	Intermolecular Forces	Dissociate	Electrophile
Bond Polarity	Hydrogen bond	Salt	Mechanism
Molecular Polarity	London Dispersion	Hydrocarbon	Arrow Pushing
Dipole	Dipole-Dipole	Solubility	Chemical Reaction
Electronegativity	Ion-Dipole	Alcohol	Physical Change

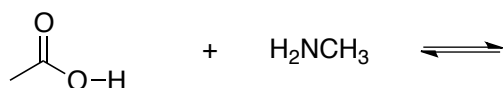
ACID-BASE CHEMISTRY*Rulez to Live By*

	Bronsted-Lowry (BL)	Lewis
Acid		
Base		

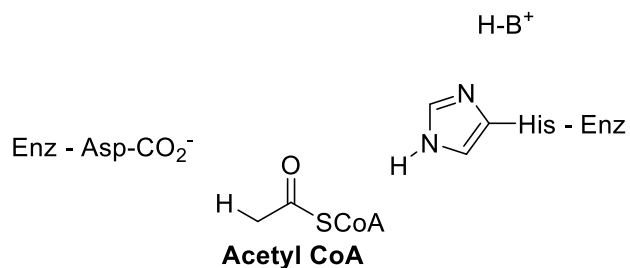
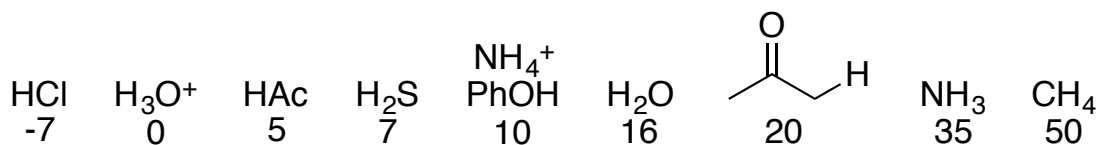
- **pKa = - log Ka = AFFINITY OF AN ACID FOR ITS PROTON**

- Equilibrium (eq \rightleftharpoons m) favors the weaker acid

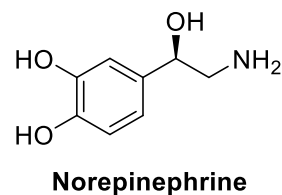
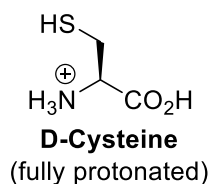
Acid Dissociation Equation: HA \rightarrow $K_a =$ $pK_a =$ *What does it mean if $K_a > 1$? If $K_a < 1$?**If $K_a = 1$?***ARROW-PUSHING** = the language / symbology of this class!*What do arrows push? What are the possible outcomes of arrow pushing?***ELECTRON RICH TO ELECTRON POOR***Identify the acid, base, conjugate base, conj. acid, and direction of equilibrium (eq \rightleftharpoons m)**Draw proper Lewis structures & mechanism for the reaction above*

Organic Acid-Base Reactions

Amino acid residues = acids & bases in **enzyme active sites**:

pKa's to Memorize

Approximating pKa's: Into which pKa family does each compound belong?



L1 & L2 HW "due" in discussion next week (quiz directly from HW) – assignment online

Next time: Electrophilic Add'n & Substitution Reaction Mechanisms